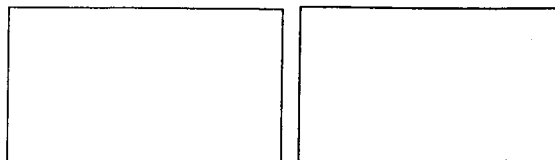
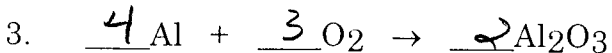
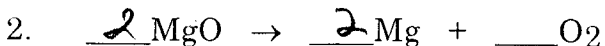
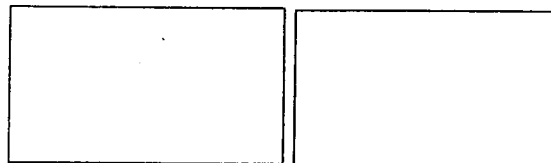
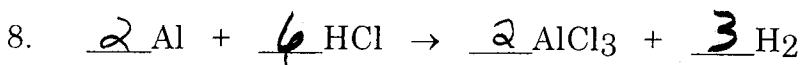


Chemistry - Unit 6 Reaction Equations Worksheet 1

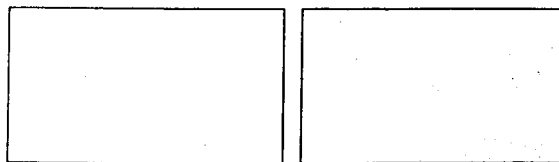
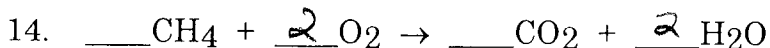
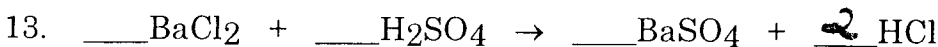
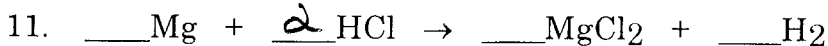
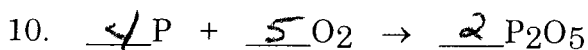
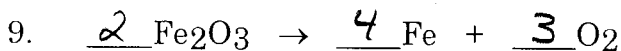
Balance the following equations by inserting the proper coefficients. For selected reactions, draw *Before* and *After* particle diagrams to show the particles involved in the reaction. Be sure to provide a key.



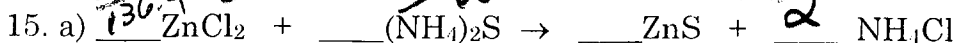
#3 *Before* *After*



#8 *Before* *After*



#14 *Before* *After*



Handwritten calculations:
65.4
35.3
136.4

Handwritten calculations:
28
32.1
68.1

b) Find the molar mass of these reactants.

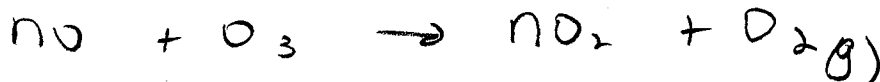
c) How many moles of ZnCl₂ would be in 25 g? How much mass would 0.55 moles of (NH₄)₂S have?

Handwritten calculation: 25 / 136.4 g/m = 0.18 mole

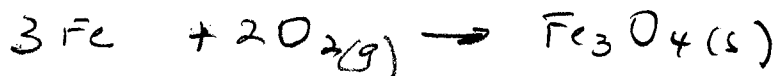
Handwritten calculation: 0.55 m x 68.1 g = 37.5 g

Part II: Write the formulas of the reactants and products, then balance the equations. (See Clues and Hints below.)

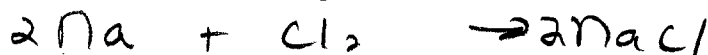
1. Nitric oxide (NO) reacts with ozone (O₃) to produce nitrogen dioxide and oxygen gas.



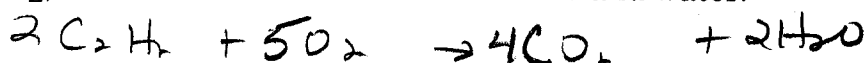
2. Iron burns in air to form a black solid, Fe₃O₄.



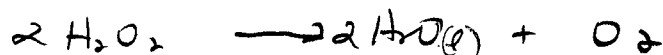
3. Sodium metal reacts with chlorine gas to form sodium chloride.



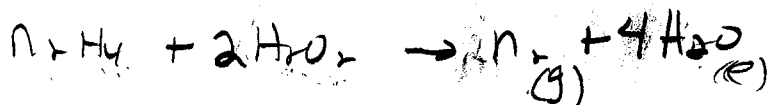
4. Acetylene, C₂H₂, burns in air to form carbon dioxide and water.



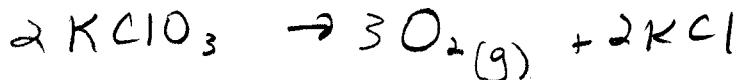
5. Hydrogen peroxide (H₂O₂) easily decomposes into water and oxygen gas.



6. Hydrazine (N₂H₄) and hydrogen peroxide are used together as rocket fuel. The products are nitrogen gas and water.



7. If potassium chlorate is strongly heated, it decomposes to yield oxygen gas and potassium chloride.



8. When sodium hydroxide is added to sulfuric acid (H₂SO₄), the products are water and sodium sulfate.



9. In the Haber process, hydrogen gas and nitrogen gas react to form ammonia, NH₃.



CLUES and HINTS:

- Products usually follow words like *produces*, *yields*, *forms*
- Watch for our diatomic elements (H₂, N₂, etc...), which are often (but not always) gases
- Include 'state subscripts' behind each substance [(s), (l), (g)] when the state is given
- Remember **air** is a mixture of (primarily) two gases, O₂ and N₂. Which is most likely to participate in a reaction?
- Elemental metals exist as single, unbonded atoms. (Ex: formula for copper metal is **Cu**)
- Watch for **ionic** vs **molecular** compounds. Use *nomenclature rules*, and your *ion chart* and *periodic table* to figure out the formulas for these.