

Chemistry: Unit 4 - Worksheet 4

More Empirical Formulas

IONS		FORMULA	NAME
1. Na^+	and Br^-	<u>NaBr</u>	<u>sodium bromide</u>
2. Cu^+	and SO_4^{2-}	<u>Cu_2SO_4</u>	<u>Copper (I) sulfate</u>
3. Pb^{2+}	and Cl^-	<u>PbCl_2</u>	<u>Lead (II) chloride</u>
4. K^+	and S^{2-}	<u>K_2S</u>	<u>potassium sulfide</u>
5. Sn^{2+}	and F^-	<u>SnF_2</u>	<u>Tin (II) fluoride</u>
6. <u>Ba^{2+}</u>	and <u>I^-</u>	BaI_2	<u>Barium Iodide</u>
7. <u>Al^{3+}</u>	and <u>Cl^-</u>	AlCl_3	<u>Aluminum Chloride</u>
8. <u>Mg^{2+}</u>	and <u>NO_3^-</u>	$\text{Mg}(\text{NO}_3)_2$	<u>Magnesium nitrate</u>
9. <u>K^+</u>	and <u>OH^-</u>	KOH	<u>potassium hydroxide</u>
10. <u>NH_4^+</u>	and <u>SO_4^{2-}</u>	$(\text{NH}_4)_2\text{SO}_4$	<u>Ammonium sulfate</u>
11. <u>Ag^+</u>	and <u>O^{2-}</u>	<u>Ag_2O</u>	silver oxide
12. <u>Li^+</u>	and <u>Br^-</u>	<u>LiBr</u>	lithium bromide
13. <u>Cu^{2+}</u>	and <u>NO_3^-</u>	<u>$\text{Cu}(\text{NO}_3)_2$</u>	copper (II) nitrate
14. <u>Mg^{2+}</u>	and <u>Cl^-</u>	<u>MgCl_2</u>	magnesium chloride
15. <u>Ca^{2+}</u>	and <u>CO_3^{2-}</u>	<u>CaCO_3</u>	calcium carbonate
16. Mg^{2+}	and NO_3^-	<u>$\text{Mg}(\text{NO}_3)_2$</u>	<u>magnesium nitrate</u>
17. Cu^{2+}	and OH^-	<u>$\text{Cu}(\text{OH})_2$</u>	<u>Copper (II) hydroxide</u>
18. <u>Na^+</u>	and <u>HCO_3^-</u>	NaHCO_3	<u>sodium bicarbonate</u>
19. <u>Fe^{3+}</u>	and <u>S^{2-}</u>	<u>Fe_2S_3</u>	iron (III) sulfide or hydrogen carbonate
20. <u>K^+</u>	and <u>CrO_4^{2-}</u>	<u>K_2CrO_4</u>	potassium chromate

Part II

Write the names of the following compounds

1. $\text{Cu}(\text{NO}_3)_2$ Copper (II) Nitrate
2. BaCl_2 Barium chloride
3. HgO Mercury (II) oxide
4. $\text{Ni}(\text{OH})_2$ Nickel (II) hydroxide
5. Na_3PO_4 Sodium phosphate
6. CaCO_3 Calcium carbonate
7. CS_2 Carbon disulfide
8. SnBr_4 Tin (IV) Bromide
9. $(\text{NH}_4)_2\text{CrO}_4$ Ammonium Chromate
10. $\text{Mg}(\text{NO}_3)_2$ Magnesium nitrate
11. Li_2O Lithium oxide
12. FeS Iron (II) Sulfide
13. N_2 Nitrogen trioxide
14. H_2SO_4 Hydrogen sulfate - sulfuric acid
15. $\text{K}_2\text{C}_2\text{O}_4$ potassium Oxalate

Part III

Write the formulas for the following compounds

1. copper (II) sulfate CuSO_4
2. sodium chromate Na_2CrO_4
3. iron (III) chloride FeCl_3
4. silver sulfide Ag_2S
5. aluminum oxide Al_2O_3
6. zinc nitrate $\text{Zn}(\text{NO}_3)_2$
7. potassium phosphate K_3PO_4
8. strontium fluoride SrF_2
9. ammonium carbonate $(\text{NH}_4)_2\text{CO}_3$
10. magnesium hydroxide $\text{Mg}(\text{OH})_2$
11. carbon tetrachloride CCl_4
12. phosphorus tribromide PBr_3
13. sulfur hexafluoride SF_6
14. sulfur dioxide SO_2
15. chromium (III) oxide Cr_2O_3
16. nitric acid HNO_3
17. hydrochloric acid HCl
18. lead(II) iodide PbI_2
19. ammonium nitrite NH_4NO_2
20. potassium hydrogen carbonate KHCO_3