

Name _____

Date _____ Pd _____

Chemistry: Unit 4 - Worksheet 4

More Empirical Formulas

IONS			FORMULA	NAME
1.	Na^+	and Br^-	_____	_____
2.	Cu^+	and SO_4^{2-}	_____	_____
3.	Pb^{2+}	and Cl^-	_____	_____
4.	K^+	and S^{2-}	_____	_____
5.	Sn^{2+}	and F^-	_____	_____
6.	_____	and _____	BaI_2	_____
7.	_____	and _____	AlCl_3	_____
8.	_____	and _____	$\text{Mg}(\text{NO}_3)_2$	_____
9.	_____	and _____	KOH	_____
10.	_____	and _____	$(\text{NH}_4)_2\text{SO}_4$	_____
11.	_____	and _____	_____	silver oxide
12.	_____	and _____	_____	lithium bromide
13.	_____	and _____	_____	copper (II) nitrate
14.	_____	and _____	_____	magnesium chloride
15.	_____	and _____	_____	calcium carbonate
16.	Mg^{2+}	and NO_3^-	_____	_____
17.	Cu^{2+}	and OH^-	_____	_____
18.	_____	and _____	NaHCO_3	_____
19.	_____	and _____	_____	iron (III) sulfide
20.	_____	and _____	_____	potassium chromate

Part II

Write the names of the following compounds

1. $\text{Cu}(\text{NO}_3)_2$ _____
2. BaCl_2 _____
3. HgO _____
4. $\text{Ni}(\text{OH})_2$ _____
5. Na_3PO_4 _____
6. CaCO_3 _____
7. CS_2 _____
8. SnBr_4 _____
9. $(\text{NH}_4)_2\text{CrO}_4$ _____
10. $\text{Mg}(\text{NO}_3)_2$ _____
11. Li_2O _____
12. FeS _____
13. NI_3 _____
14. H_2SO_4 _____
15. $\text{K}_2\text{C}_2\text{O}_4$ _____

Part III

Write the formulas for the following compounds

- | | |
|--------------------------------|--|
| 1. copper (II) sulfate _____ | 2. sodium chromate _____ |
| 3. iron (III) chloride _____ | 4. silver sulfide _____ |
| 5. aluminum oxide _____ | 6. zinc nitrate _____ |
| 7. potassium phosphate _____ | 8. strontium fluoride _____ |
| 9. ammonium carbonate _____ | 10. magnesium hydroxide _____ |
| 11. carbon tetrachloride _____ | 12. phosphorus tribromide _____ |
| 13. sulfur hexafluoride _____ | 14. sulfur dioxide _____ |
| 15. chromium (III) oxide _____ | 16. nitric acid _____ |
| 17. hydrochloric acid _____ | 18. lead(II) iodide _____ |
| 19. ammonium nitrite _____ | 20. potassium hydrogen carbonate _____ |