

## Unit 4 – Worksheet 2

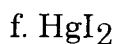
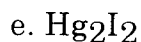
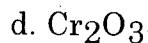
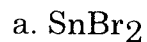
### Ionic Compounds

#### Properties

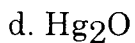
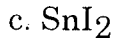
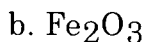
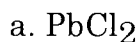
#### Basic structural unit

1. Give the name of the following simple binary ionic compounds.
  - a.  $\text{Na}_2\text{O}$
  - b.  $\text{K}_2\text{S}$
  - c.  $\text{MgCl}_2$
  - d.  $\text{CaBr}_2$
  - e.  $\text{BaI}_2$
  - f.  $\text{Al}_2\text{S}_3$
  - g.  $\text{CsBr}$
  - h.  $\text{AgF}$
2. Give the name of the following simple binary ionic compounds.
  - a.  $\text{Na}_3\text{N}$
  - b.  $\text{K}_2\text{O}$
  - c.  $\text{AgBr}$
  - d.  $\text{MgI}_2$
  - e.  $\text{SrO}$
3. Write the formula for the following binary ionic compounds.
  - a. lithium bromide
  - b. sodium iodide
  - c. silver sulfide
  - d. cesium oxide
  - d. beryllium iodide
  - f. barium hydride
  - g. aluminum fluoride
  - h. potassium oxide
4. Write the formula for these ionic substances.
  - a. silver oxide
  - b. aluminum sulfide
  - c. sodium nitride
  - d. barium chloride
  - e. strontium hydride

5. Write the name of these ionic substances using a Roman numeral to specify the charge of the cation.



6. Write the name of these ionic substances using a Roman numeral to specify the charge of the cation.



7. Write the formulas of each ionic compound.

a. chromium (III) chloride

b. tin (IV) oxide

c. lead (II) oxide

d. copper (II) iodide

e. cobalt (II) oxide

f. cobalt (III) oxide

8. Write the formulas of each ionic compound.

a. chromium (III) sulfide

b. manganese (IV) oxide

c. gold (III) chloride

d. titanium (IV) chloride

e. iron (II) bromide

f. iron (II) oxide