

SMALL-SCALE LAB: Names and Formulas of Ionic Compounds**LABORATORY RECORD SHEET**

Use with SECTION 6.4

SAFETY

Wear safety glasses and follow the standard safety procedures, as outlined on page 5 of this manual.

PURPOSE

To observe the formation of compounds, and to write their names and formulas.

MATERIALS

- pencil
- paper
- reaction surface
- chemicals shown in Figure A
- ruler

PROCEDURE

On a separate sheet of paper, draw a grid similar to Figure A. Make each square 2 cm on each side. Draw black Xs in each square of the grid. Place a reaction surface over the grid and add the chemicals as shown in Figure A. Use the grid in Figure A as a data table to record your observations for each solution.

	AgNO ₃ (Ag ⁺)	Pb(NO ₃) ₂ (Pb ²⁺)	CaCl ₂ (Ca ²⁺)
Na ₂ CO ₃ (CO ₃ ²⁻)	a	e	i
Na ₃ PO ₄ (PO ₄ ³⁻)	b	f	j
NaOH (OH ⁻)	c	g	k
Na ₂ SO ₄ (SO ₄ ²⁻)	d	h	l

Figure A

ANALYSIS

Using the experimental data, record the answers to the following questions.

1. Describe each precipitate that forms as milky, grainy, cloudy, or gelatinous. Which mixture(s) did not form a precipitate?

2. Write the formulas and names of the chemical compounds produced in the mixings.

YOU'RE THE CHEMIST



Use the space below to write your observations to the small-scale activities in the *You're the Chemist* section.