

Periodic Table Hopscotch Laboratory:

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Students' Guide

Goals

- Visualize the concept of the mole
- Gain experience in calculating grams and moles

The Activity

In this activity, students visualize the concept of the mole using a mole of chalk as a model. They practice calculations of moles and grams, and end up with a better understanding of what a mole is and how chemists use it.

Materials for Each Group

- Data book and pencil
- A chunk of chalk
- Semi-analytical balance
- A table of atomic masses
- A calculator

SAFETY

No special safety considerations are required.

Instructions

1. Take a chunk of chalk and measure its mass on the balance: _____
2. Go outside and draw something on the sidewalk, but **DO NOT USE THE ENTIRE PIECE OF CHALK.**
3. Go back in class and weigh the unused chalk _____
4. Based on the initial mass of the chalk and the mass at the end, calculate how many grams of chalk you left out on the side walk: _____
5. Write down the molecular formula of the chalk: _____
6. Get the masses of the atoms: _____
7. Calculate the chalk's molar mass: _____
8. Using this data calculate:
 - a. How many moles of chalk did you leave on the sidewalk? _____
 - b. How many molecules of chalk did you leave on the sidewalk? _____
 - c. How many atoms of calcium did you use? _____
 - d. How many atoms of carbon did you use? _____
 - e. How many atoms of oxygen did you use? _____

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Summary

How much does a mole of chalk weigh? _____

How many grams did you leave outside? _____

Is it (circle the correct answer):

1. less than a mole
2. equal to a mole
3. more than a mole

Define a mole, and explain its importance: _____
