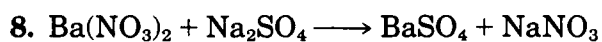
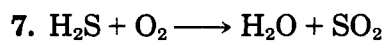
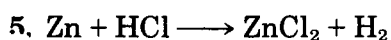
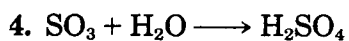
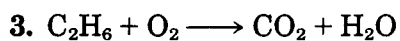
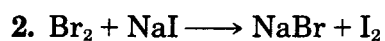
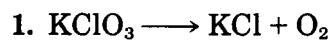


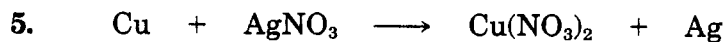
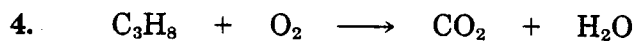
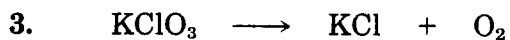
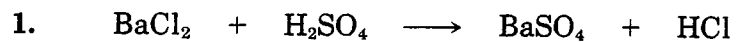
Balancing Equations

Balance the following equations using coefficients. Remember that you must not change symbols or formulas in order to balance an equation.



A Balancing Act

Balance each of the following equations. You may wish to refer to the rules of balancing in Figure 8-7 on page 187 of your textbook.



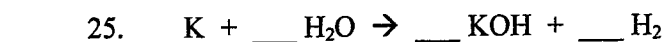
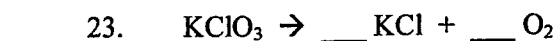
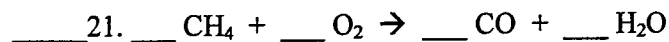
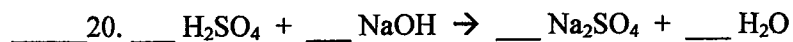
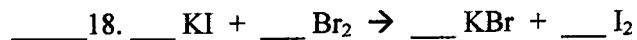
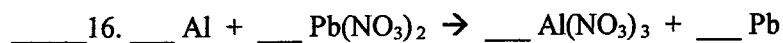
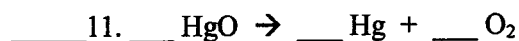
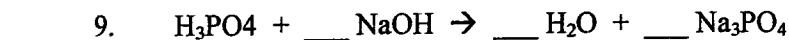
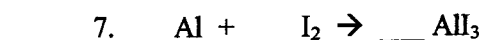
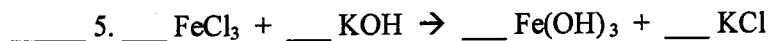
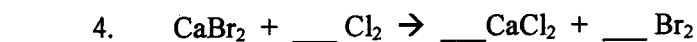
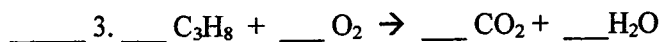
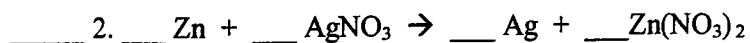
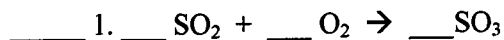
CHEMICAL EQUATIONS

NAME _____

BALANCE the following EQUATIONS:

1. $\text{___N}_2 + \text{___O}_2 \longrightarrow \text{___N}_2\text{O}_5$
2. $\text{___NaClO}_3 \longrightarrow \text{___NaCl} + \text{___O}_2$
3. $\text{___Al} + \text{___Cl}_2 \longrightarrow \text{___AlCl}_3$
4. $\text{___Pb(NO}_3)_2 \longrightarrow \text{___PbO} + \text{___NO}_2 + \text{___O}_2$
5. $\text{___K} + \text{___H}_2\text{O} \longrightarrow \text{___KOH} + \text{___H}_2$
6. $\text{___FeS}_2 + \text{___O}_2 \longrightarrow \text{___Fe}_2\text{O}_3 + \text{___SO}_2$
7. $\text{___C}_3\text{H}_7\text{OH} + \text{___O}_2 \longrightarrow \text{___CO}_2 + \text{___H}_2\text{O}$
8. $\text{___ZnCl}_2 + \text{___KOH} \longrightarrow \text{___Zn(OH)}_2 + \text{___KCl}$
9. $\text{___Na}_2\text{CO}_3 + \text{___HCl} \longrightarrow \text{___NaCl} + \text{___CO}_2 + \text{___H}_2\text{O}$
10. $\text{___Al} + \text{___Pb(NO}_3)_2 \longrightarrow \text{___Al(NO}_3)_3 + \text{___Pb}$
11. $\text{___SiO}_2 + \text{___C} \longrightarrow \text{___SiC} + \text{___CO}$
12. $\text{___Cu} + \text{___Cl}_2 \longrightarrow \text{___CuCl}$
13. $\text{___P} + \text{___O}_2 \longrightarrow \text{___P}_4\text{O}_{10}$
14. $\text{___C} + \text{___H}_2\text{O} \longrightarrow \text{___CO} + \text{___H}_2$
15. $\text{___Zn} + \text{___H}_2\text{SO}_4 \longrightarrow \text{___ZnSO}_4 + \text{___H}_2$
16. $\text{___Ba(OH)}_2 + \text{___H}_2\text{SO}_4 \longrightarrow \text{___BaSO}_4 + \text{___H}_2\text{O}$
17. $\text{___CO} + \text{___O}_2 \longrightarrow \text{___CO}_2$
18. $\text{___KI} + \text{___Cl}_2 \longrightarrow \text{___KCl} + \text{___I}_2$
19. $\text{___Cu} + \text{___AgNO}_3 \longrightarrow \text{___Cu(NO}_3)_2 + \text{___Ag}$
20. $\text{___ZnSO}_4 + \text{___(NH}_4)_2\text{S} \longrightarrow \text{___ZnS} + \text{___(NH}_4)_2\text{SO}_4$

Balance the equations for the following reactions. Classify each reaction as one of the following types of reactions: C-Combination; D-Decomposition; SR-Single Replacement; DR-Double Replacement; CC-Complete Combustion; IC-Incomplete Combustion



ix. Balance each of the following. If already balanced, write BAL.

List type of rxn

Type of rxn

