

Chemistry – Unit 1 - Worksheet 6

Dimensional Analysis

Use the factor-label method to make the following conversions. Remember to use the appropriate number of sf's in your answer.

Part 1

1. $74 \text{ cm} \times \frac{1 \text{ m}}{10^{+2} \text{ cm}} = 7.4 \times 10^{-6} \text{ meters}$
2. $8.32 \times 10^{-2} \text{ kg} \times \frac{10^3 \text{ g}}{1 \text{ kg}} = 8.32 \times 10^1 \text{ grams}$
3. $55.5 \text{ mL} \times \frac{1 \text{ cm}^3}{1 \text{ mL}} = 55.5 \text{ cm}^3$
4. $0.00527 \text{ cal} \times \frac{1 \text{ kcal}}{10^3 \text{ cal}} = 5.27 \times 10^{-6} \text{ kilocalories}$
5. $9.52 \times 10^{-4} \text{ m} \times \frac{10^6 \text{ } \mu\text{m}}{1 \text{ m}} = 9.52 \times 10^2 \text{ micrometers}$
6. $41.0 \text{ mL} \times \frac{1 \text{ L}}{10^3 \text{ mL}} = 4.10 \times 10^{-2} \text{ liters}$
7. $6.0 \times 10^{-1} \text{ g} \times \frac{10^3 \text{ mg}}{1 \text{ g}} = 6.0 \times 10^2 \text{ mg}$
8. $8.34 \times 10^{-9} \text{ cg} \times \frac{1 \text{ g}}{10^2 \text{ cg}} = 8.34 \times 10^{-11} \text{ g}$
9. $5.0 \times 10^3 \text{ mm} \times \frac{1 \text{ m}}{10^3 \text{ mm}} = 5.0 \text{ m}$
10. $1 \text{ day} \times \frac{24 \text{ hrs}}{1 \text{ day}} \times \frac{60 \text{ min}}{1 \text{ hr}} \times \frac{60 \text{ sec}}{1 \text{ min}} = 8.64 \times 10^4 \text{ seconds}$
11. $5 \times 10^4 \text{ mm} \times \frac{1 \text{ m}}{10^3 \text{ mm}} \times \frac{1 \text{ km}}{10^3 \text{ m}} = 5 \times 10^{-2} \text{ km}$
12. $9.1 \times 10^{-13} \text{ kg} \times \frac{10^3 \text{ g}}{1 \text{ kg}} \times \frac{10^9 \text{ ng}}{1 \text{ g}} = 9.1 \times 10^{-1} \text{ ng}$
13. $1 \text{ year} \times \frac{365 \text{ days}}{\text{yr}} \times \frac{24 \text{ hr}}{1 \text{ day}} = 8.76 \times 10^3 \text{ hours (approximately)}$

$$14. \quad 4.22 \text{ cL} \times \frac{1 \text{ L}}{10^2 \text{ cL}} \times \frac{10^3 \text{ mL}}{1 \text{ L}} = 4.22 \times 10^1 \text{ mL}$$

$$15. \quad 1 \text{ mile} \times \frac{5280 \text{ ft}}{1 \text{ mi}} \times \frac{12 \text{ in}}{1 \text{ ft}} = 6.336 \times 10^4 \text{ inches}$$

Part 2

1. How many nickels could you trade for 250 yen? \$1 = 150 yen.

$$250 \text{ yen} \times \frac{\$1}{150 \text{ yen}} \times \frac{20 \text{ nickels}}{1 \$} \sim 3.3 \times 10^1 \text{ or } 33 \text{ nickels}$$

2. Your school club sold 600 tickets to a chili supper. The chili recipe for 10 persons requires 2 teaspoons of chili powder? How many teaspoons of chili powder will you need altogether?

$$600 \text{ tickets} \times \frac{1 \text{ person}}{1 \text{ ticket}} \times \frac{2 \text{ teaspoons}}{10 \text{ people}} = 1.2 \times 10^2 \text{ tsp}$$

3. How many cups of chili powder will you need? Three teaspoons (tsp) equal one tablespoon (TBS) and 16 tablespoons equal 1 cup.

$$1.2 \times 10^2 \text{ tsp} \times \frac{1 \text{ T}}{3 \text{ tsp}} \times \frac{1 \text{ cup}}{16 \text{ T}} = 2.5 \text{ cups}$$

4. How many seconds in a year? (assume 30 days in an average month)

$$3.1536 \times 10^7 \text{ sec} \quad 1 \text{ year} \times \frac{365 \text{ days}}{\text{year}} \times \frac{24 \text{ hrs}}{1 \text{ day}} \times \frac{60 \text{ min}}{\text{hr}} \times \frac{60 \text{ sec}}{1 \text{ min}}$$

5. Chloroform is a liquid once used for anesthetic. What is the volume of 5.0 g of chloroform.

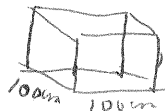
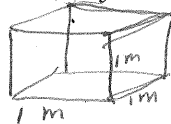
The density of chloroform 1.49 g/mL

$$5.0 \text{ g} \times \frac{1 \text{ mL}}{1.49 \text{ g}} = 3.4 \text{ mL}$$

6. How many inches long is a football field?

$$100 \text{ yds} \times \frac{3 \text{ ft}}{\text{yd}} \times \frac{12 \text{ in}}{\text{ft}} = 3.6 \times 10^3 \text{ in}$$

7. How many m³ is 4.6 cm³? Express your answer in scientific notation.



$$1 \text{ m}^3 = 1 \times 10^6 \text{ cm}^3$$

$$4.6 \text{ cm}^3 \times \frac{1 \text{ m}^3}{1 \times 10^6 \text{ cm}^3}$$

8. How many mg is 59.0 kg? Express your answer in scientific notation.

$$59.0 \text{ kg} \times \frac{10^3 \text{ g}}{1 \text{ kg}} \times \frac{10^3 \text{ mg}}{1 \text{ g}}$$

$$4.6 \times 10^{-6} \text{ m}^3$$

$$5.90 \times 10^7 \text{ mg}$$