

CHAPTER 18 REVIEW ACTIVITY

Text Reference: ~~Pg. 200-207~~

Solution Vocabulary

Choose words from the list to fill in the blanks in the paragraphs.
A term may be used more than once.

Word List

- | | |
|-----------------------|-------------------------|
| alloy | size |
| amalgam | solubility |
| aqueous solution | solubility curve |
| concentrated solution | solute |
| dilute solution | solution |
| effervescence | solvent |
| heterogeneous mixture | specific gravity |
| homogeneous mixture | stirring |
| molality | supersaturated solution |
| molarity | temperature |
| pressure | tincture |
| saturated solution | unsaturated solution |

A physical combination of substances in which the components are uniformly mixed is called a(n) (1). Such a combination in which the components are individual molecules, atoms, or ions is called a(n) (2).

If two substances in such a combination were originally in different phases, the substance that changed phase is said to be dissolved in the other and is called the (3). The substance that did not change phase and did the dissolving is called the (4).

Such a combination in which water does the dissolving is called a(n) (5). One in which alcohol does the dissolving is called a(n) (6). Such a combination of two or more metals is called a(n) (7); if one of the metals is mercury, the combination is called a(n) (8).

A measure of the amount of one substance that can be dissolved in another is called its (9) in that other substance. A factor that, when increased, generally causes this quantity to increase for solids and decrease for gases is (10). A factor that, when increased, increases this quantity for gases is (11). The rapid escape of a dissolved gas from a liquid is called (12).

During the process of solution, certain factors influence the rate of solution formation. These include the (13) of the solute particles, the (14) of the solvent-solute mixture, the rate of (15), or agitation, of the mixture, and the amount of (16) already dissolved.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____
11. _____
12. _____
13. _____
14. _____
15. _____
16. _____

Name _____

REVIEW ACTIVITY Chapter 16

Solution Vocabulary (continued)

The amount of a substance that dissolves in another substance for varying values of some factor such as temperature is shown in a(n) (17). When one substance is dissolved as much as possible in another substance, under given conditions, the combination is called a(n) (18). When more of the dissolved substance can be dissolved, the combination is called a(n) (19). If the combination is unstable and has more of the dissolved substance than it usually can hold under the given conditions, it is called a(n) (20). A combination in which a small amount of a substance is dissolved is called a(n) (21). One in which a large amount is dissolved is called a(n) (22).

The ratio of the mass of a substance to the mass of an equal volume of water is called the (23) of the substance. The number of moles of a substance dissolved in one cubic decimeter of a solution is called the (24) of the solution. The number of moles of a substance dissolved in one kilogram of another substance is called the (25) of the solution.

- 17. _____
- 18. _____
- 19. _____
- 20. _____
- 21. _____
- 22. _____
- 23. _____
- 24. _____
- 25. _____